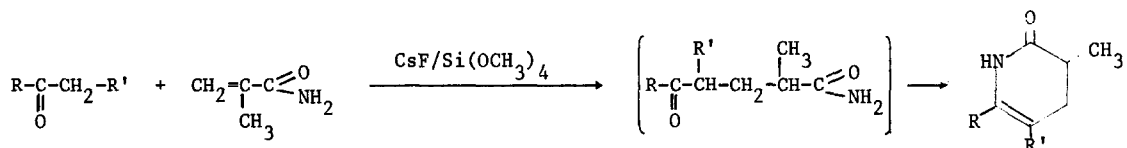
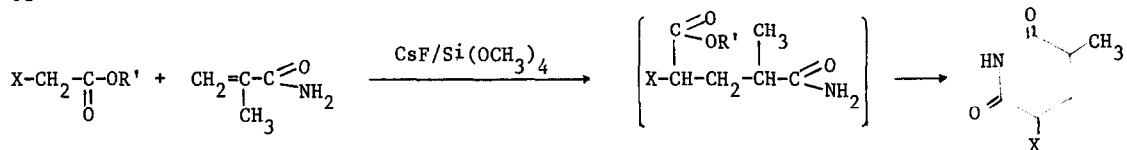


concerning the 1,4-addition of ketones and β -cyano or β -keto esters to methacrylamide when $\text{CsF}/\text{Si}(\text{OCH}_3)_4$ was used in the same conditions as before. In this case, the Michael type reaction occurs and is followed, in situ, by the cyclisation, leading to 3,4-dihydro 2(1H)-pyridinones or 3,5-disubstituted glutarimides.



or



The typical procedure is the following one : synthesis of 3-methyl 5,6-diphenyl 3,4-dihydro 2(1H)-pyridinone.

25 mmol of methacrylamide $\text{CH}_2=\overset{\text{O}}{\parallel}{\text{C}}-\underset{\text{CH}_3}{\text{C}}-\text{NH}_2$, 25 mmol of phenylacetophenone are added to

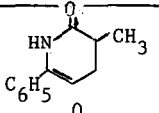
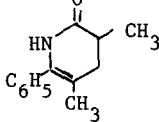
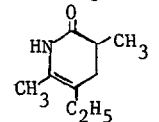
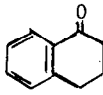
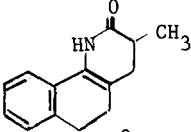
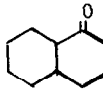
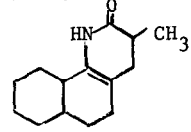
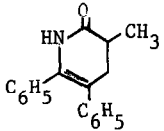
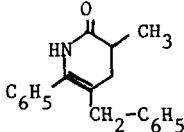
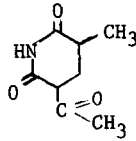
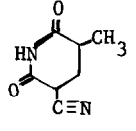
15 mmol of $\text{Si}(\text{OCH}_3)_4$ and 20 mmol of CsF under nitrogen atmosphere.

The mixture was stirred and heated at 80°C for 5 hours. Hydrolysis was not necessary and the mixture was extracted with CH_2Cl_2 and the crude product was recrystallized from ethyl acetate (m p : 157°C).

Even, if some ketones (cyclohexanone, methyl isopropyl ketone, 2,4-pentanedione) give mixtures, in many cases this method allows the formation of the pure cyclic compound in a one pot process, without hydrolysis and in good yields while stepwise reactions are usually necessary (10).

The mechanism we propose is the one previously indicated : nucleophilic activation of $\text{Si}(\text{OCH}_3)_4$ by the fluoride ion giving a basic species which promotes enolate formation. The fast silylation of this enolate gives the silyl enol ether which promotes the 1,4 adduct on the α,β -unsaturated primary amide. This compound reacts in situ with the alcohol present in the mixture to give the 1,5-difunctional compound. This primary amide undergoes a fast cyclisation in situ and the final product is isolated by recrystallisation.

Michael additions on methacrylamide in the presence of CsF/Si(OCH₃)₄

Electrophile	Michael donor	Reaction conditions		Isolated product	Yield(%)
		t(h)	T°C		
$\text{CH}_2=\text{C}(\text{CH}_3)-\text{C}(=\text{O})\text{NH}_2$	$\text{C}_6\text{H}_5-\text{C}(=\text{O})-\text{CH}_3$	2	80		(a) 55
"	$\text{C}_6\text{H}_5-\text{CH}_2-\text{C}(=\text{O})-\text{CH}_3$	12	80		(a) 46
"	$\text{CH}_3-\text{C}(=\text{O})-\text{CH}_2-\text{CH}_2-\text{CH}_3$	12	80		33
"		12	100		94
"		12	100		90
"	$\text{C}_6\text{H}_5-\text{CH}_2-\text{C}(=\text{O})-\text{C}_6\text{H}_5$	5	80		76
"	$\text{C}_6\text{H}_5-\text{CH}_2-\text{C}(=\text{O})-\text{CH}_2-\text{C}_6\text{H}_5$	6	80		70
"	$\text{CH}_3-\text{C}(=\text{O})-\text{CH}_2-\text{C}(=\text{O})-\text{OC}_2\text{H}_5$	2	80		(b) 83
"	$\text{N}\equiv\text{C}-\text{CH}_2-\text{C}(=\text{O})-\text{OC}_2\text{H}_5$	12	80		(c) 84

(a) see ref. 13 (b) see ref. 12 (c) see ref. 11.

Elemental analysis and spectral data (IR and NMR) are consistent with the structures of the products.

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